

HOW TO MAKE HOT ICE

Overview and objective:

Students will discover how to **make hot or sodium** acetate through crystallization.

Crystallization is also a chemical solid–liquid separation technique, in which mass transfer of a solute from the liquid solution to a pure solid crystalline phase occurs.

Source: [Wikipedia.org/wiki/Crystallization](https://www.wikipedia.org/wiki/Crystallization)

Presented by

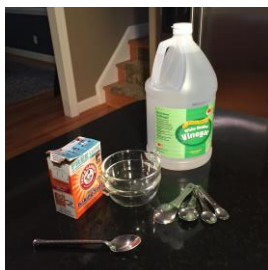
BASF
We create chemistry

WHAT YOU WILL NEED:

- Baking soda
- White vinegar
- Stovetop or burner
- Spoon
- Heat-safe container or saucepan
- Two small glass bowls
- One large glass bowl
- Ice
- Water
- Safety goggles
- Safety gloves
- Adult supervision

Procedure:

1. Create homemade sodium acetate by mixing six tablespoons of baking soda with two liters of white vinegar. A chemical reaction will occur.
2. Once fizzy reaction stops, boil mixture on medium-high heat until about 90% of liquid is evaporated and a crusty film begins to form on top.
3. As crystals form when liquid evaporates, scrape them off sides of pan with a spoon and add to a small glass bowl. Set aside for later.
4. Remove completed solution from heat and pour into small, glass dish. Add a tablespoon of vinegar and stir.
5. Fill larger glass bowl with ice and water add smaller glass filled with solution into ice water until chilled.
6. Once solution is chilled, take a pinch of crystallized sodium acetate and add to solution to create hot ice.



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